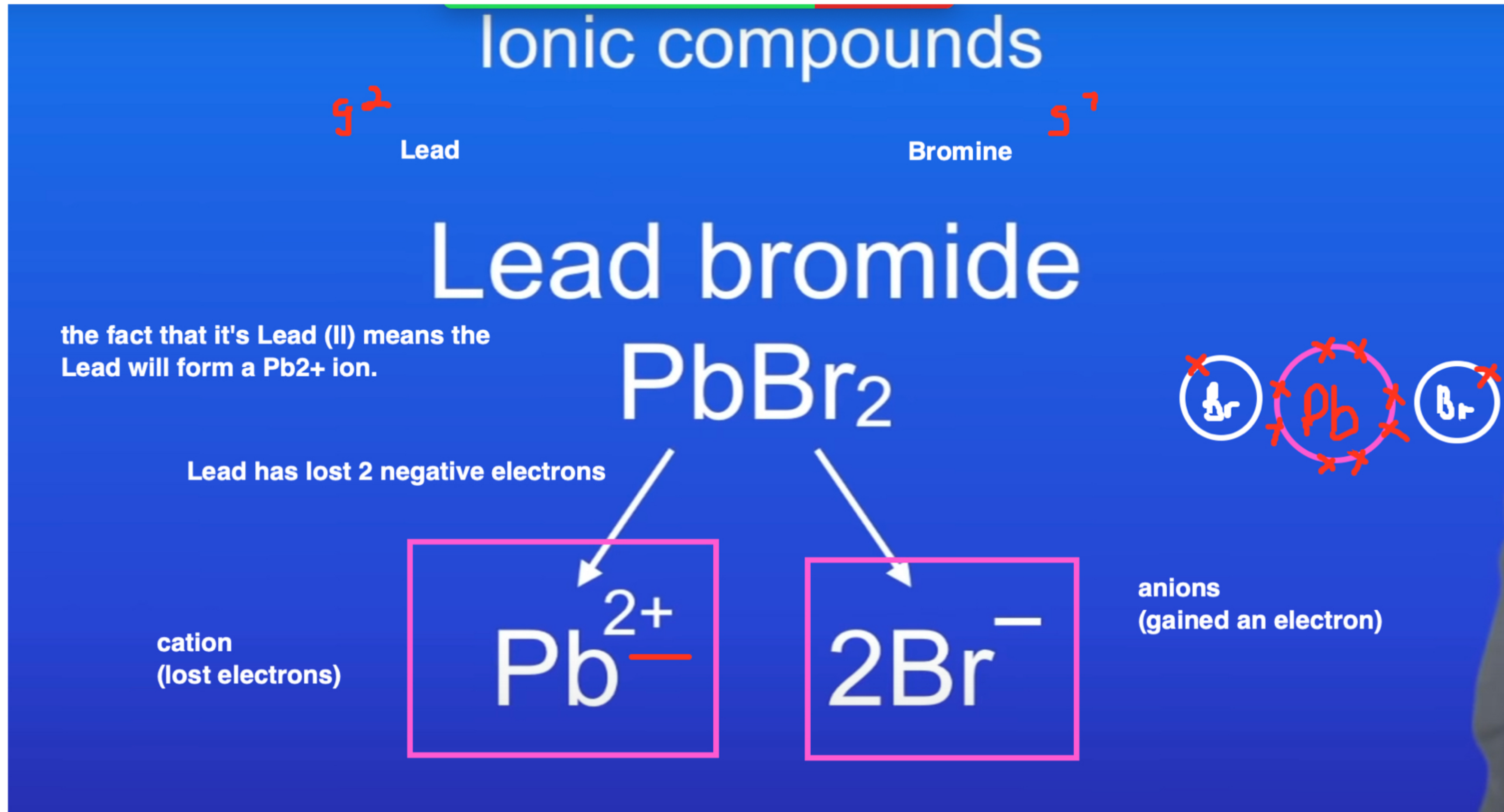


What is Electrolysis?

Electrolysis is a process in which electrical energy, from a direct current (dc) supply, breaks down electrolytes. The free moving ions in electrolytes are attracted to the oppositely charged electrodes which connect to the dc supply.



Molten Electrolysis

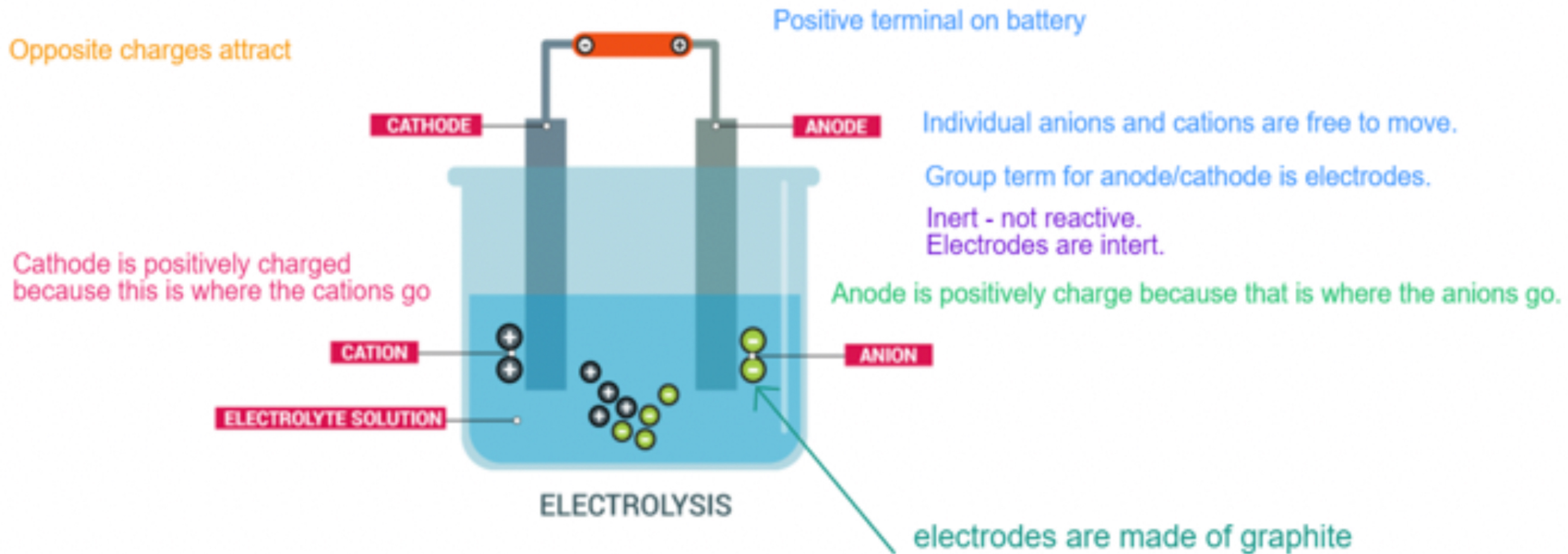
Ionic compounds can conduct electricity when molten or in solution.

This is because the ions are free to move, and carry the charge. Whereas when they are solid they are not free to move.

If we melt an ionic substance it can conduct electricity.

If we apply a potential difference (voltage) then current will flow and we can make a circuit.

Opposite charges attract

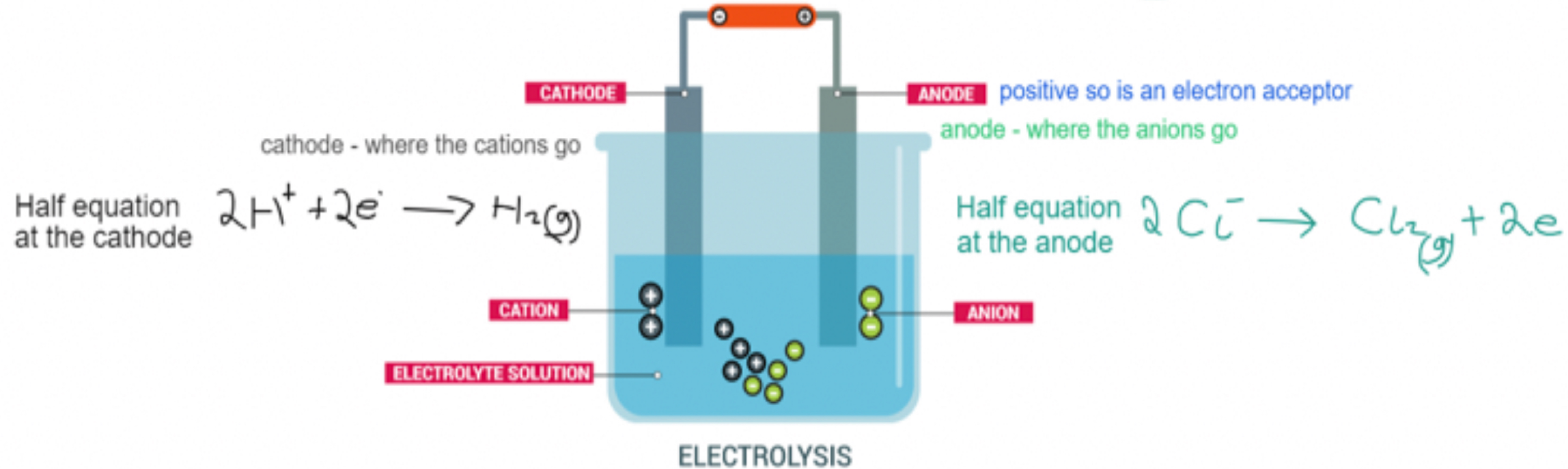


Electrolysis

Ionic compounds can only conduct electricity when molten or aqueous (dissolved in water).

As these are the only forms where the ions are free to move.

The Electrolysis Of Molten Hydrogen Chloride (HCl)



Anode and cathode are collectively known as electrodes

